

## REVIEW

# Plasma-based liquefaction of methane: The road from hydrogen production to direct methane liquefaction

Ramses Snoeckx<sup>1,2</sup> | Alexander Rabinovich<sup>2</sup> | Danil Dobrynin<sup>2</sup> |  
Annemie Bogaerts<sup>1</sup> | Alexander Fridman<sup>2</sup>

<sup>1</sup> Research Group PLASMANT, Department of Chemistry, University of Antwerp, Antwerp, Belgium

<sup>2</sup> AJ Drexel Plasma Institute, Drexel University, Philadelphia, Pennsylvania, USA

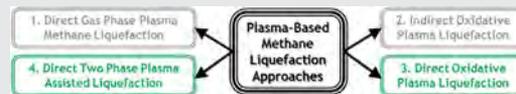
**Correspondence**

Ramses Snoeckx, Research Group PLASMANT, Department of Chemistry, University of Antwerp, Universiteitsplein 1, Antwerp BE-2610, Belgium.  
Email: ramses.snoeckx@uantwerpen.be

**Funding information**

Advanced Plasma Solutions; Drexel University; Belgian Federal Office for Science Policy (BELSPO); Fund for Scientific Research Flanders, Grant numbers: G038316N, V403616N

For the energy industry, a process that is able to transform methane—being the prime component of natural gas—efficiently into a liquid product would be equivalent to a goose with golden eggs. As such it is no surprise that research efforts in this field already date back to the nineteen hundreds. Plasma technology can be considered to be a novel player in this field, but nevertheless one with great potential. Over the past decades this technology has evolved from sole hydrogen production, over indirect methane liquefaction to eventually direct plasma-assisted methane liquefaction processes. An overview of this evolution and these processes is presented, from which it becomes clear that the near future probably lies with the direct two phase plasma-assisted methane liquefaction and the far future with the direct oxidative methane liquefaction.

**KEYWORDS**

direct methane incorporation, dry reforming of methane, methane liquefaction, non-thermal plasma, partial oxidation of methane

## 1 | INTRODUCTION

### 1.1 | Methane liquefaction

Interest for transforming methane into more valuable liquid products dates back as far as the beginning of the 20th century. At present, several promising methane sources are available, such as the regular natural gas, biogas, landfill gas, methanated CO<sub>2</sub>, and shale gas, all coming with their own advantages and drawbacks depending on the viewpoint one takes. Although some sources may be debatable, one cannot neglect the fact that to date methane is one of the cheapest sources of energy available in the world.<sup>[1]</sup> As such it is not surprising that a very large fraction is used for heating and electrical power generation, for which methane in many respects is an attractive fuel—besides its current price point. Nevertheless, it can be considered to be one of the most

underutilized feedstocks with regard to the production of value-added chemicals and liquid fuels, the main obstacle being its physical condition, the—industrially undesirable—gaseous state. As such, one can imagine that a lot of research effort has already been performed for the development of efficient natural gas conversion technologies. The most well-developed are the indirect synthesis routes, which first convert methane into the intermediate syngas, a mixture of CO and H<sub>2</sub>, with the main processes being steam methane reforming (SMR), dry reforming of methane (DRM), and partial oxidation of methane (POX). The syngas is then further processed by Methanol or Fischer-Tropsch synthesis to obtain the desired liquid products.<sup>[2]</sup> Although, successful from a technical perspective, these methods are marked by low overall yields and high energy inputs.<sup>[3]</sup> Their direct (thermo-catalytic) synthesis counterparts circumvent the expensive and energy intensive syngas step, but are

technologically still very challenging and costly, while achieving no better results concerning product yields.<sup>[2]</sup>

## 1.2 | Plasma technology

Unsurprisingly, interest for alternative (non-conventional) reforming processes took a spike and one of the alternatives considered to have great potential in this area is plasma technology.<sup>[4,5]</sup> The advantage of (non-thermal) plasmas is that the gas can be “activated” by electron impact excitation, ionization, and dissociation reactions, instead of the need for heating the entire reactor. Several types of plasmas have already been used for the conversion of methane,<sup>[6–9]</sup> including dielectric barrier discharges (DBD),<sup>[7–39]</sup> microwave discharges,<sup>[7–9,25,40–44]</sup> glow discharges,<sup>[7,9,25,45–48]</sup> coronas,<sup>[7–9,25,46,49–51]</sup> sparks,<sup>[9,22,25,38,52,53]</sup> gliding arcs,<sup>[6–9,22,25,54–61]</sup> radio frequency (RF) plasmas,<sup>[9,25,62,63]</sup> and thermal plasmas.<sup>[5,7,9,64–68]</sup>

Figure 1 gives an overview of the different directions that are being investigated regarding plasma-based methane liquefaction approaches. The “direct gas phase plasma methane liquefaction,” is the conversion of pure methane into hydrogen gas and liquid hydrocarbon chains originating from the remaining CH<sub>2</sub> blocks. The “oxidative plasma liquefaction” methods, can be subdivided in an indirect and a direct approach. The indirect approach converts methane into syngas together with an oxidant, such as O<sub>2</sub>, CO<sub>2</sub>, or H<sub>2</sub>O, known as partial oxidation of methane (POX), dry reforming of methane (DRM), and steam methane reforming (SMR), respectively. The syngas is then further processed into liquids using Fischer-Tropsch or methanol synthesis. The direct approach, on the other hand, tries to convert methane with the same oxidants into oxygenated liquid products, such as alcohols and aldehydes, in one step. Finally, recently a new approach “direct two phase plasma-assisted liquefaction” has emerged, which—as its name suggests—aspire the direct liquefaction of methane, through its incorporation into a second phase, namely existing liquid hydrocarbons.

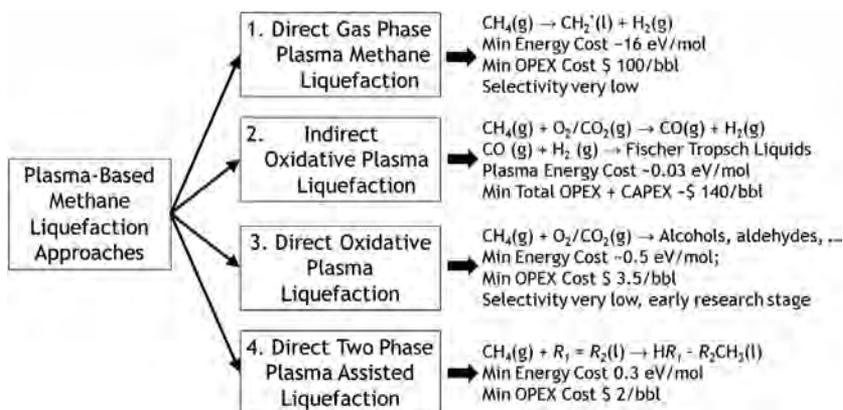


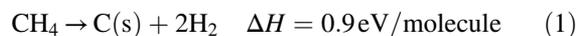
FIGURE 1 Overview of the different plasma-based methane liquefaction approaches

This review will give an overview of the efforts made to date in the field of plasma-based liquefaction of methane. Looking at the minimal energy cost in eV/molecule—and subsequently the resultant operational expenditure (OPEX) cost—for these different approaches in Figure 1, it already becomes clear that to date the two most promising techniques are the “direct oxidative plasma liquefaction” and the “direct two phase plasma-assisted liquefaction.” The former still needs a lot of research and development regarding an increase of the selectivity, before it can be successfully applied on an industrial scale, and thus becomes important in the long run, while the latter, if successfully up scaled, can already be implemented on a short term.

## 2 | DIRECT GAS PHASE PLASMA METHANE LIQUEFACTION

The direct conversion of methane in the absence of oxidants such as O<sub>2</sub>, CO<sub>2</sub>, and H<sub>2</sub>O, better known as “the pyrolysis of methane” is a technique which can be used to synthesize higher hydrocarbons.<sup>[25]</sup>

Total pyrolysis:



Partial pyrolysis:



Under equilibrium conditions (i.e., thermal plasma arc discharge) the synthesis of H<sub>2</sub> and C<sub>2</sub>H<sub>2</sub> (acetylene) is already being successfully applied on an industrial scale. It reaches a conversion of 95–98% and an acetylene selectivity of 90–95%, with an energy consumption of 2.5–3 eV/molecule (see also Table 1).<sup>[69]</sup> The same process for atmospheric pressure non-thermal plasma discharges, on the other hand, has a much higher energy consumption, varying from 14 to 25 eV/molecule in a pulsed spark discharge, 17–21 eV/molecule for a streamer discharge, 38–57 eV/molecule in a pulsed DC DBD and a staggering 116–175 eV/molecule for an AC DBD.<sup>[38]</sup> Table 1 gives an overview of the results for these different plasma discharges.

One might ask why there is still such a big interest for these non-thermal approaches. The answer is: not only do they have the advantage of operating at room temperature, in comparison to the thermal process which yields acetylene and H<sub>2</sub>, but they can also yield a wide variety of products with a higher commercial value, including: ethane, propane, butane, ethylene, acetylene, isobutene, and hydrogen.<sup>[21,22]</sup> Furthermore, it has been

TABLE 1 Overview of methane pyrolysis results

Discharge type	Conversion (%)	Energy consumption (eV/molecule)	Refs.
Gliding arc	34	0.82	[70]
Hollow cathode	26	1.19	[22]
Rotating arc	42	1.56	[22]
Spark AC	83	2.21	[22]
Arc DC (thermal)	95	3.0	[69]
Spark	70	6.1	[25]
Corona	25	52	[25]
DBD AC	15	82.9	[22]
DBD AC	5–8	116–175	[38]
Pulsed microwave	90–100	1.0	[71]
Pulsed double pointed	23.5	3.8	[25]
Pulsed spark	49	6.94	[22]
Pulsed spark	18–69	14–25	[38]
Pulsed electron beam	5–7	18	[25]
Pulsed streamer	19–41	17–21	[38]
Pulsed DBD	9	19	[22]
Pulsed DBD DC	6–13	38–57	[38]
Gliding arc + catalyst (Al <sub>2</sub> O <sub>3</sub> )	18	2.4–7.3	[25]
Spark + catalyst (Fe, Co, Ni on HZSM-5)	40	19	[72]
DBD + catalyst (SiO <sub>2</sub> )	45	52	[25]

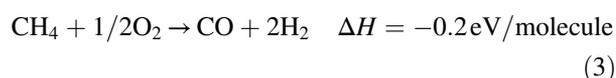
shown that the plasma characteristics, that is, electron energy and the degree of “warmness,” greatly affects the product selectivity. Plasma sources that have strong non-thermal characteristics show higher C<sub>2</sub>H<sub>6</sub> selectivity, while, upon increasing degree of warmness, the selectivity of H<sub>2</sub> and C<sub>2</sub>H<sub>2</sub> increases. As such the proper selection of a plasma source can be used as a tool to control the reaction pathway.<sup>[22]</sup>

Nevertheless, looking at Figure 1 and Table 1 it becomes clear that non-thermal plasma discharges are inefficient for the (partial) pyrolysis of methane, because of the high activation energy required by electron impact processes.<sup>[24]</sup> The combination of a catalyst together with the reactive plasma species might be the key to enhancing the selectivity, and therefore the energy efficiency to targeted products. Indeed, numerical simulations show that vibrationally excited methane species are the abundant long-lived excited species. Due to the fact that these species are highly reactive on transition metal catalysts, a synergistic effect between DBD and Ni-catalysts can be achieved.<sup>[24]</sup> These vibrationally excited methane states will also be important for the last technique described in this review, that is, “the direct two phase plasma-assisted liquefaction.”

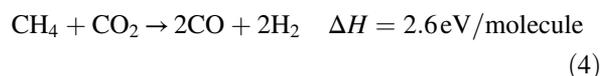
### 3 | OXIDATIVE PLASMA LIQUEFACTION

Although plasma technology is based on a different principle than the conventional thermal methods, pure methane reforming can have the same drawback of carbon deposition due to total pyrolysis (reaction 1). In other applications this might be beneficial like for the pure production of hydrogen, carbon nanotubes and carbon black.<sup>[5,73,74]</sup> However, in the present case every carbon molecule deposited as soot is a loss of building blocks and eventually energy. Furthermore, the deposited carbon tends to be very conductive, which, depending on the process, can influence the plasma discharge(s).<sup>[75]</sup> As such, in analogy with their conventional counter parts, an oxidizing agent was added to minimize this carbon formation and deposition,<sup>[76]</sup> the candidates being again O<sub>2</sub>, CO<sub>2</sub>, and H<sub>2</sub>O yielding partial oxidation (reaction 3), dry reforming of methane (reaction 4) and steam methane reforming (reaction 5), respectively. These processes yield mainly the syngas components CO and H<sub>2</sub>, which can be further processed into Fischer-Tropsch liquids or methanol. An additional reason for adding an oxidizing agent is the aspiration for the direct formation of value-added oxygenated products such as formaldehyde, methanol, ethanol, dimethyl ether, formic acid, etc. These two options are therefore called the “indirect and direct oxidative plasma liquefaction” approaches, respectively. In both areas a lot of research effort is also directed to the combination of a plasma system with a catalyst, which falls outside the scope of the current review and for which we refer to the work of Neyts et al.<sup>[73]</sup>

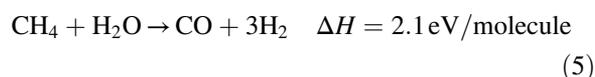
Partial oxidation of methane:



Dry reforming of methane:



Steam methane reforming:



#### 3.1 | Indirect oxidative plasma liquefaction

The “indirect oxidative plasma liquefaction” approach consists of producing liquid products from methane in two stages. The first stage entails the successful production of the syngas components CO and H<sub>2</sub>, by plasma-based partial oxidation, dry reforming of methane, or steam methane reforming. Subsequently, this syngas is then further processed into liquid chemicals and fuels using the traditional

catalytic Fischer-Tropsch processes or methanol synthesis. Although industrially possible, this two-step process has a high capital cost, since in its essence, we are oxidizing the carbon atoms present in the first step, only to reduce them again afterwards in the second step. This also explains the high—combined OPEX and APEX—cost as presented in Figure 1, despite its extremely low energy cost (in the case of partial oxidation of methane, see below).

### 3.1.1 | Partial oxidation of methane

The reaction enthalpy shows that partial oxidation of methane is a slightly exothermic process. Taking this into account, the process can be realized with a very low energy input, by selective generation of radicals, charged and excited species, stimulating the low temperature long chain reactions.<sup>[9]</sup> (Near-)atmospheric pressure non-equilibrium discharges have proven to be effective for this purpose. The different plasma systems include pulsed corona discharges,<sup>[9,50]</sup> microwave discharges,<sup>[8,9,25,41,44]</sup> dielectric barrier discharges,<sup>[8,9,13,25–29]</sup> and gliding arc discharges.<sup>[8,9,25,58,59]</sup> For lower pressure non-equilibrium discharges, results have been obtained for radio frequency,<sup>[9,63]</sup> and pulsed discharges.<sup>[9]</sup> In general, the non-thermal discharges can obtain high energy efficiencies, but fail when it comes to the industrially requested high production (flow) rates of the syngas. One of the exceptions are the gliding arc discharges, which can generate a strongly non-equilibrium, mainly non-thermal plasma, even at relatively high power levels.<sup>[77]</sup> To date the best results regarding a combination of yield and energy efficiency have been achieved with a reverse vortex flow gliding arc discharge,<sup>[9]</sup> as shown in Table 2.

### 3.1.2 | Dry reforming of methane

Dry reforming of methane has the advantage of using carbon dioxide as oxidizing agent, thus allowing the utilization of the main greenhouse gas as a feedstock and leading to a “greener” reforming of methane, especially when the syngas is intended for further conversion.<sup>[11]</sup> However, the thermodynamics for dry reforming of methane (DRM) are not as feasible as for the partial oxidation of methane (POX). As a result the thermodynamic barrier makes processes relying on heat unfavorable. As such it is no surprise that plasma technology is amongst the most popular of the unconventional approaches reported in literature.<sup>[11]</sup> Plasmas allow for a high conversion of methane and many of them show positive results regarding the energy balance.<sup>[11]</sup> Just like for the partial oxidation, several kinds of non-thermal plasmas have been applied for DRM, such as corona discharge,<sup>[7,8,25,51]</sup> dielectric barrier discharge (DBD),<sup>[7,11–14,17–20,25,30–37]</sup> microwave discharge,<sup>[7,40]</sup> atmospheric pressure glow discharge (APGD),<sup>[7,46]</sup> spark discharge,<sup>[53]</sup> and gliding arc discharge.<sup>[7,8,25,55,60,61]</sup> In general, the application of non-equilibrium plasma allows for dry

**TABLE 2** Overview of oxidative plasma liquefaction for partial oxidation of methane

Discharge type	Conversion (%)	Energy consumption (eV/molecule)	Refs.
Microwave (thermal)	100	0.5	[25]
Arc DC (thermal)	100	1	[78]
Vortex rotating arc	30	1.4–5.0	[79]
Rotating arc	78–99	1.6–2.2	[80]
DBD 75 kHz	14–27	5–13	[25]
DBD 30 kHz	24	13	[26]
DBD micro reactor	45	17	[81]
DBD + catalyst (Yt + Zr)	50	88	[25]
DBD + catalyst (CZA)	29–45	89–159	[82]

reforming of methane at a relatively low temperature without using any conventional catalyst. Although, depending on the plasma source, the energy efficiency can exceed 40%, the total consumption of electric energy can be relatively high in the process.<sup>[9]</sup> The type of plasma which is currently most often used for investigating DRM (and gas conversion in general) is the dielectric barrier discharge (DBD).<sup>[12,76]</sup> However, when evaluating the existing non-thermal technologies, it appears that arcs, together with microwave and glow discharges, show the best performances because of their high energetic densities and their ability to create a large reactive volume.<sup>[5,8,9]</sup>

In an extensive review by Tao et al.<sup>[7]</sup> regarding dry reforming of methane by different plasmas, it was concluded that three key factors need to be taken into account for achieving high conversions and energy efficiency, that is, the electron density, plasma temperature and reactor configuration. As such, the treatment capacity for corona discharge and DBD are restricted due to their non-uniformity, low electron density and limited reaction volume, while microwave discharges, gliding arc discharges and APGDs possess high electron energy and electron density, as well as the proper plasma temperature. Table 3 shows an overview of the obtained results in the field of “indirect oxidative plasma liquefaction” research for dry reforming of methane.

### 3.1.3 | Steam methane reforming

A third option to be considered as oxidizing agent is H<sub>2</sub>O, leading to the so-called steam methane reforming (SMR). A lot of research is still ongoing in this area. In general, the same kind of non-thermal plasmas can be, and have been, used for SMR as for DRM,<sup>[25,91–95]</sup> and the same principles hold due to the endothermic nature of both processes. The main

**TABLE 3** Overview of oxidative plasma liquefaction for dry reforming of methane

Discharge type	Conversion (%)	Energy consumption (eV/molecule)	Refs.
Pulsed corona	16	0.7	[25]
Glow	54	0.89	[45]
Arc DC (thermal)	~8	1	[65]
Cold plasma jet	38.7	2.09	[83]
Microwave	70	3.18	[84]
Glow	50.5	3.24	[85]
Gliding arc	9.8	3.44	[60]
Radio frequency	61.9	4.44	[86]
Gliding Arc	35.5	6.30	[55]
Corona DC	83.3	11.8	[87]
Corona	55.1	18.6	[46]
Gliding arc AC	8	20	[25]
DBD (0.4–2 kHz)	9.8–22.8	38–70	[25]
DBD (2–40 kHz)	25–65	73–100	[25]
DBD (30 kHz)	35–75	52–85	[25]
DBD	61.3	84.2	[88]
DBD	58.6	75.6	[37]
Cold plasma jet + catalyst (NiAl <sub>2</sub> O <sub>3</sub> )	48.2	1.39	[83]
DBD + catalyst (starch)	48	16	[25]
Corona + catalyst (Zeolite)	39.6	41.7	[89]
DBD + catalyst (NiAl <sub>2</sub> O <sub>3</sub> )	44.6	107.6	[90]

difference is that SMR can generate a maximum hydrogen yield and/or high syngas ratio, since both CH<sub>4</sub> and H<sub>2</sub>O are converted into H<sub>2</sub>, while the presence of O from H<sub>2</sub>O prevents soot formation. The use of H<sub>2</sub>O compared to O<sub>2</sub> or CO<sub>2</sub> as oxidizing agent allows several options to deliver the H<sub>2</sub>O in the plasma, e.g., as liquid, vapor or steam. Depending on this (energetic) state of the H<sub>2</sub>O, the electrical characteristics, plasma chemistry, conversions and energy efficiency vary widely even within the same discharge type. A detailed analysis of the latter is outside the scope of this review paper, but a brief summary of results obtained in literature for plasma-assisted steam methane reforming is shown in Table 4.

### 3.2 | Direct oxidative plasma liquefaction

In contrast to the indirect approach, the “direct oxidative plasma liquefaction” approach tries to synthesize oxygenated liquid products in one step. The main product categories that can be defined are: alcohols (e.g., methanol, ethanol),

**TABLE 4** Overview of oxidative plasma liquefaction for steam methane reforming

Discharge type	Conversion (%)	Energy consumption (eV/molecule)	Refs.
Gliding arc	48–71	1.08–1.26	[91]
Spark	55	5.3	[25]
MW	91.6	28	[92]
Corona AC	28–80	33–105	[93]
DBD	17	53	[94]
DBD	50	54	[95]
Gliding arc + catalyst (Ni/Al <sub>2</sub> O <sub>3</sub> )	73–82	0.90–1.33	[91]
Spark + catalyst (Ni)	70	1.6	[25]
DBD + catalyst (Ni/SiO <sub>2</sub> )	66–86	3.8–18.4	[95]

aldehydes (e.g., formaldehyde), ethers (e.g., dimethyl ether), and acids (e.g., formic acid, oxalic acid). To achieve this, in general the same techniques and setups are used as for the indirect approach, but the operating conditions are tuned in order to try to produce these more valuable products in higher amounts compared to the aforementioned syngas. In general, DRM favors the formation of H<sub>2</sub>, CH<sub>2</sub>O, CH<sub>3</sub>CHO, and CH<sub>2</sub>CO, while POX yields higher concentrations of H<sub>2</sub>O<sub>2</sub>, CH<sub>3</sub>OH, C<sub>2</sub>H<sub>5</sub>OH, CH<sub>3</sub>OOH, and C<sub>2</sub>H<sub>5</sub>OOH, as revealed by computer simulations.<sup>[13]</sup> In both processes CO is formed and during POX significant amounts of undesired CO<sub>2</sub> can be formed. An explanation for this difference in product formation is proposed and is attributed to the crucial role played by the CH<sub>3</sub> radicals.<sup>[13]</sup> During DRM these radicals primarily recombine to hydrocarbons, while during POX the formation of CH<sub>3</sub>O<sub>2</sub> radicals is favored. The production of these CH<sub>3</sub>O<sub>2</sub> radicals lead to the formation of methanol, and subsequently formaldehyde.<sup>[13]</sup> Two other computational studies using a zero-dimensional chemical kinetics model obtained the same findings for DRM and also illustrated the tradeoff between conversion and energy efficiency.<sup>[11,12]</sup> As a result it was posed that classical DBDs would probably not be competitive, and that it would be more interesting from a combined economic and ecological point of view to focus on more efficient discharges, such as microwave discharges, gliding arcs, and packed bed DBDs.<sup>[12]</sup> Furthermore, notwithstanding all the research in this area, the main products formed currently remain the syngas components CO and H<sub>2</sub> and the gaseous C<sub>2</sub> hydrocarbons, ethane, ethylene, and acetylene, with few studies reporting the production of significant amounts of liquid hydrocarbons.<sup>[1,7,11,76]</sup> Recent efforts are therefore, among others, focused on the synergism between plasma and catalysis for this process, in which the underlying idea is for the plasma to activate the inert gas

molecules and to have the catalyst to react and/or recombine these reactive species in a selective way towards the desired (liquid) products.<sup>[73,74]</sup> From these findings it becomes clear that although the “direct oxidative liquefaction of methane” is one of the most promising in theory, a lot of research is still needed to improve the yield of the desired (oxygenated) liquid products.

#### 4 | DIRECT TWO PHASE PLASMA-ASSISTED LIQUEFACTION

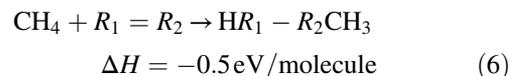
Reports show that the API gravity of extracted crude oil has been decreasing,<sup>[96]</sup> this means that the most useful and valuable part, that is, the light crude, is decreasing. Combined with the limited supplies of fossil hydrocarbons, attention is being focused on new technologies to process the cheaper heavy crude fractions. The aim is to convert these heavier crudes into more economic valuable products by reforming, cracking or upgrading.<sup>[97–100]</sup> The most common methods used in refineries are thermal and catalytic cracking. All these traditional methods are carried out at relatively high temperatures and moderate pressures, which often results in the production of high amounts of coke as unwanted byproduct. In this field as well, plasma technology can be used as a way to convert a wide variety of hydrocarbons. In such a plasma cracking process, the overall reactions are the same as for conventional cracking. However, energy and free radicals provided by the plasma are used for the cracking—instead of heat—allowing the plasma cracking process to operate at ambient temperature and atmospheric pressure unlike the traditional non-catalytic cracking processes.<sup>[100]</sup>

Several promising results have already been achieved for the cracking of heavier crudes using plasma technology, including pure liquid plasma reformers and combinations with active gases such as methane and ethane mixtures with N<sub>2</sub>, He, and Ar.<sup>[96,98–100]</sup> In this cracking process the long chain hydrocarbons are split into shorter chain hydrocarbons, which then react with the radicals in the gas phase to form a termination product. Although promising, a lot of energy is still being put in the endothermic process of the hydrocarbon bond cleavage and radical production. An energetic more favorable approach would therefore be preferred.

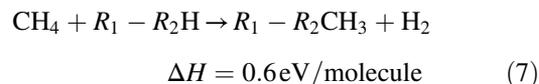
This knowledge of plasma based hydrocarbon cracking, in combination with the mixed results obtained when trying to achieve the liquefaction of methane through the above mentioned reforming paths (mainly just resulting in the production of syngas), led to another novel approach. This new approach can be seen as a combination of the above mentioned plasma based reforming and cracking processes, namely, the direct

liquefaction of methane by its incorporation in existing liquid hydrocarbons. Two possible pathways for the successful incorporation are suggested<sup>[101,102]</sup>:

Saturation of double bonds:



Polymerization:



The saturation of double bonds (reaction 6) is the preferred process, since it is exothermic with an eventual plasma energy cost of only 0.3 eV per molecule of CH<sub>4</sub>, due to its exothermicity, while all the other hydrocarbon reactions in non-thermal plasmas, such as polymerization (reaction 7) and dissociation, require much higher energy inputs, mainly due to their endothermic character, making them ineffective for the energy efficient direct liquefaction of CH<sub>4</sub>. To promote the saturation process (and hence its energy efficiency), local heating needs to be avoided, by working in the non-equilibrium discharge region.

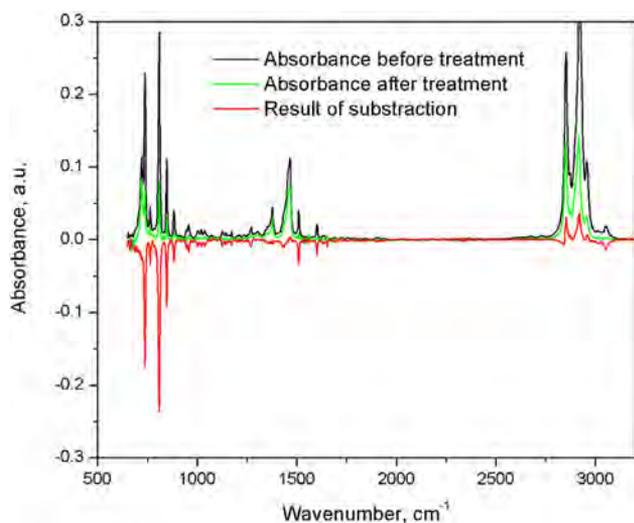
The key to this approach lies in the selective vibrational excitation of CH<sub>4</sub>, which allows for an exothermic plasma catalytic incorporation with extremely low energy cost.<sup>[102]</sup> Theoretically, this allows the energy cost to remain beneath 0.3 eV/molecule, which would be a significant improvement over conventional gas-to-liquid Fischer-Tropsch synthesis.<sup>[102]</sup> The only plasma systems known to be able to provide the necessary vibrational-translational non-equilibrium at atmospheric pressure are high gas flow gliding arc, microwave, and atmospheric pressure glow discharges, which have a vibrational temperature of 2000–4000 K. Another requirement for the effective incorporation at low energy cost (0.3–0.5 eV/molecule) is a gas temperature below 700–900 K in the reactor.<sup>[102]</sup>

For two plasma systems the successful incorporation of CH<sub>4</sub> into liquid hydrocarbons at this low energy cost has already been achieved: an atmospheric pressure nanosecond pulsed DBD and an atmospheric pressure glow discharge. Both discharges were ignited inside gaseous bubbles fed through liquid hydrocarbons; this bubbling is used in order to increase the reaction surface area as well as the mixing.<sup>[101–103]</sup> In the DBD system a continuous flow of 0.2 SLPM was applied and gas samples were taken after the reactor. While for the APGD system a fixed volume of gas (i.e., 2.97 L) was continuously recycled and a gas sample was taken after 5 min. Table 5, shows the gas composition before and after plasma treatment. In both cases the decrease in CH<sub>4</sub> volume could not be explained

**TABLE 5** Gas composition before and after DBD and APGD treatment of a CH<sub>4</sub>/N<sub>2</sub> mixture in diesel<sup>[101]</sup>

	DBD treatment		APGD treatment	
	Without discharge	Treatment (continuous)	Without discharge	Treatment (batch-recycle)
N <sub>2</sub> (L)	0.106	0.106	0.27	0.27
CH <sub>4</sub> (L)	0.094	0.093	2.7	2.313
C <sub>2</sub> H <sub>6</sub> (L)	$3 \times 10^{-5}$	$17 \times 10^{-5}$	$19 \times 10^{-5}$	$21 \times 10^{-5}$
C <sub>2</sub> H <sub>2</sub> (L)	Not detectable	$3 \times 10^{-5}$	Not detectable	$8 \times 10^{-5}$
C <sub>2</sub> H <sub>4</sub> (L)	Not detectable	Not detectable	Not detectable	Not detectable
H <sub>2</sub> (L)	Not detectable	$4 \times 10^{-4}$	Not detectable	$54 \times 10^{-5}$
Total (L)	0.200	0.199	2.97	2.583
Difference (L)	0.001 L		0.387 L	

by the formation of its dissociation products C<sub>2</sub>H<sub>6</sub>, C<sub>2</sub>H<sub>2</sub>, and H<sub>2</sub>. Leaving the CH<sub>4</sub> incorporation into liquid as the only reasonable explanation. This thesis was supported by NMR and FTIR analysis of the treated liquid samples combined with the GC analysis of the gas phase, which showed the effective fixation of methane in the liquid fuel, a corresponding decrease of the methane concentration in the gas phase, as well as structural changes to the liquid hydrocarbons (i.e., double bond saturation and aromatic ring opening).<sup>[101]</sup> Figure 2 shows the FTIR spectra of the liquid sample before and after plasma treatment and its difference, significant changes can be seen for C–H bond absorption at 2850–3100 cm<sup>-1</sup> and phenyl functional groups at 700–800 cm<sup>-1</sup>. The increase of saturated C–H bond and decrease in phenyl functional groups and unsaturated C–H bonds suggests that phenyl rings are opened, and the activated methane saturates the carbon double bonds. Although this proves that the concept is feasible with non-thermal plasmas, further research is targeted toward a better understanding of the

**FIGURE 2** FTIR spectrum of the liquid sample before and after plasma treatment with methane<sup>[101]</sup>

chemical processes taking place and increasing the conversion rates.

## 5 | CONCLUSION

An overview of the progress in the field on plasma based liquefaction of methane has been given. From all the available literature, ranging from experiments, to model predictions and physicochemical insights, we can draw the following conclusions. It becomes clear that plasma technology can definitely play an important role in this field and it is not beyond our grasp to think about the eventual industrial implementation of this technology. Nevertheless, as always only a few candidates seem suitable for the specific task at hand. To be more precise, we can identify two candidates, each for a more specific task and with a different timeline on the implementation horizon.

In the long run it seems evident that the “direct oxidative liquefaction” has the highest potential for the efficient production of value-added liquid chemicals, such as alcohols, aldehydes, esters, and acids. Furthermore, from both experimental and computational results it appears that partial oxidation of methane is more suited than dry reforming of methane, although the use of CO<sub>2</sub> presents a more ecologically attractive oxidant. For now, however, a lot of research in this area is still needed to increase the selectivity towards these valuable bulk chemicals over the currently produced syngas, possibly by combinations with catalysts. Regarding the treatment capacity and hence the industrial applicability, corona discharges and DBDs seem restricted, while microwave, gliding arc and atmospheric pressure glow discharges possess the necessary properties, i.e., high electron energy and electron density, as well as proper plasma temperature.

In the shorter run the “direct two phase plasma-assisted liquefaction” could already deliver on the basic promise of liquefying methane. The latter is achievable not by its complete transformation into value-added products, such as is

the case for the “direct oxidative liquefaction,” but rather by increasing the value of to date unusable—and hence invaluable—crudes by its successful incorporation, based on double bond saturation and polymerization reactions with vibrationally excited methane.

From the current analysis it becomes evident that the later solution can reach industrial implementation at a faster pace. Nevertheless, the former path should not be abandoned, since—when successful—it offers us the possibility of producing a wide variety of value-added chemicals and fuels starting from the same building block and allowing the flexibility to tune the output depending on the market's needs.

## ACKNOWLEDGMENTS

This work was in part supported by Advanced Plasma Solutions (APS) and Drexel University. The authors acknowledge financial support from the IAP/7 (Interuniversity Attraction Pole) program “PSI-Physical Chemistry of Plasma-Surface Interactions,” financially supported by the Belgian Federal Office for Science Policy (BELSPO), as well as the Fund for Scientific Research Flanders (FWO; grant no. G038316N and grant no. V403616N).

## REFERENCES

- [1] J.-M. Lavoie, *Front. Chem.* **2014**, *2*, 1.
- [2] J. H. Lunsford, *Catal. Today* **2000**, *63*, 165.
- [3] J. R. H. Ross, *Catal. Today* **2005**, *100*, 151.
- [4] S. Samukawa, M. Hori, S. Rauf, K. Tachibana, P. Bruggeman, G. Kroesen, J. C. Whitehead, A. B. Murphy, A. F. Gutsol, S. Starikovskaia, U. Kortshagen, J.-P. Boeuf, T. J. Sommerer, M. J. Kushner, U. Czarnetzki, N. Mason, *J. Phys. D. Appl. Phys.* **2012**, *45*, 253001.
- [5] E. Tatarova, N. Bundaleska, J. P. Sarrette, C. M. Ferreira, *Plasma Sources Sci. Technol.* **2014**, *23*, 63002.
- [6] A. Gutsol, A. Rabinovich, A. Fridman, *J. Phys. D. Appl. Phys.* **2011**, *44*, 274001.
- [7] X. Tao, M. Bai, X. Li, H. Long, S. Shang, Y. Yin, X. Dai, *Prog. Energy Combust. Sci.* **2011**, *37*, 113–124.
- [8] G. Petitpas, J. Rollier, A. Darmon, J. Gonzalez-Aguilar, R. Metkemeijer, L. Fulcheri, *Int. J. Hydrogen Energy* **2007**, *32*, 2848–2867.
- [9] A. Fridman, *Plasma Chemistry*. Cambridge University Press, New York **2008**.
- [10] R. Snoeckx, M. Setareh, R. Aerts, P. Simon, A. Maghari, A. Bogaerts, *Int. J. Hydrogen Energy* **2013**, *38*, 16098.
- [11] R. Snoeckx, R. Aerts, X. Tu, A. Bogaerts, *J. Phys. Chem. C* **2013**, *117*, 4957.
- [12] R. Snoeckx, Y. X. Zeng, X. Tu, A. Bogaerts, *RSC Adv.* **2015**, *5*, 29799.
- [13] C. De Bie, J. van Dijk, A. Bogaerts, *J. Phys. Chem. C* **2015**, *119*, 22331.
- [14] Y. Zhang, Y. Li, Y. Wang, C. Liu, B. Eliasson, *Fuel Process. Technol.* **2003**, *83*, 101.
- [15] V. Goujard, J.-M. Tatibouët, C. Batiot-Dupeyrat, *Plasma Chem. Plasma Process.* **2011**, *31*, 315.
- [16] T. Nozaki, V. Goujard, S. Yuzawa, S. Moriyama, A. Ağral, K. Okazaki, *J. Phys. D. Appl. Phys.* **2011**, *44*, 274010.
- [17] X. Tu, J. C. Whitehead, *Appl. Catal. B Environ.* **2012**, *125*, 439.
- [18] G. Scarduelli, G. Guella, D. Ascenzi, P. Tosi, *Plasma Process. Polym.* **2011**, *8*, 25.
- [19] A. Janeco, N. R. Pinha, V. Guerra, *J. Phys. Chem. C* **2015**, *119*, 109.
- [20] A. Ozkan, T. Dufour, G. Arnoult, P. De Keyzer, A. Bogaerts, F. Reniers, *J. CO2 Util.* **2015**, *9*, 74.
- [21] K. Thanyachotpaiboon, S. Chavadej, T. A. Caldwell, L. L. Lobban, R. G. Mallinson, *AIChE J.* **1998**, *44*, 2252.
- [22] D. H. Lee, Y.-H. Song, K.-T. Kim, J.-O. Lee, *Plasma Chem. Plasma Process.* **2013**, *33*, 647.
- [23] K. Konno, K. Onoe, Y. Takiguchi, T. Yamaguchi, *J. Fuels* **2014**, *2014*, 286392.
- [24] T. Nozaki, K. Okazaki, *Catal. Today* **2013**, *211*, 29.
- [25] A. I. Pushkarev, A. Zhu, X. Li, R. V. Sazonov, *High Energy Chem.* **2009**, *43*, 202.
- [26] L. Zhou, B. Xue, U. Kogelschatz, B. Eliasson, *Plasma Chem. Plasma Process.* **1998**, *18*, 375.
- [27] T. Nozaki, A. Ağral, S. Yuzawa, J. G. E. Han Gardeniers, K. Okazaki, *Chem. Eng. J.* **2011**, *166*, 288.
- [28] V. Goujard, T. Nozaki, S. Yuzawa, A. Ağral, K. Okazaki, *J. Phys. D. Appl. Phys.* **2011**, *44*, 274011.
- [29] D. W. Larkin, L. L. Lobban, R. G. Mallinson, *Catal. Today* **2001**, *71*, 199.
- [30] J. Sentek, K. Krawczyk, M. Młotek, M. Kalczyńska, T. Kroker, T. Kolb, A. Schenk, K. H. Gericke, K. Schmidt-Szałowski, *Appl. Catal. B Environ.* **2010**, *94*, 19.
- [31] Y. Li, C. J. Liu, B. Eliasson, Y. Wang, *Energy Fuels* **2002**, *16*, 864.
- [32] B. Eliasson, C. Liu, U. Kogelschatz, *Ind. Eng. Chem. Res.* **2000**, *39*, 1221.
- [33] V. Goujard, J.-M. Tatibouët, C. Batiot-Dupeyrat, *Appl. Catal. A Gen.* **2009**, *353*, 228.
- [34] X. Zhang, M. S. Cha, *J. Phys. D. Appl. Phys.* **2015**, *48*, 215201.
- [35] L. M. Martini, G. Dilecce, G. Guella, A. Maranzana, G. Tonachini, P. Tosi, *Chem. Phys. Lett.* **2014**, *593*, 55.
- [36] V. J. Rico, J. L. Hueso, J. Cotrino, A. R. González-Elipe, *J. Phys. Chem. A* **2010**, *114*, 4009.
- [37] Q. Wang, B.-H. Yan, Y. Jin, Y. Cheng, *Plasma Chem. Plasma Process.* **2009**, *29*, 217.
- [38] X. Li, A. Zhu, K. Wang, Y. Xu, Z. Song, *Catal. Today* **2004**, *98*, 617.
- [39] R. G. Mallinson, C. M. Sliepcevich, S. Rusek, *Prepr. Pap., Am. Chem. Soc., Div. Fuel Chem.* **1988**, *32*, 266.
- [40] B. Fidalgo, A. Domínguez, J. Pis, J. Menéndez, *Int. J. Hydrogen Energy* **2008**, *33*, 4337.
- [41] T.-S. Kim, S. Song, K.-M. Chun, S. H. Lee, *Energy* **2010**, *35*, 2734.

- [42] M. Jasinski, M. Dors, H. Nowakowska, G. V. Nichipor, J. Mizeraczyk, *J. Phys. D. Appl. Phys.* **2011**, *44*, 194002.
- [43] T. Kovács, R. T. Deam, *J. Phys. D. Appl. Phys.* **2006**, *39*, 2391.
- [44] A. Oumghar, J. C. Legrand, A. M. Diamy, N. Turillon, *Plasma Chem. Plasma Process.* **1995**, *15*, 87.
- [45] D. Li, X. Li, M. Bai, X. Tao, S. Shang, X. Dai, Y. Yin, *Int. J. Hydrogen Energy* **2009**, *34*, 308.
- [46] Y. Yang, *Ind. Eng. Chem. Res.* **2002**, *41*, 5918.
- [47] D. Wei, Y. Hui, C. Qi, Y. Yongxiang, D. Xiaoyan, *Plasma Sci. Technol.* **2005**, *7*, 3132.
- [48] K. Hiraoka, K. Aoyama, K. Morise, *Can. J. Chem.* **1985**, *63*, 2899.
- [49] A. Beloqui Redondo, E. Troussard, J. A. van Bokhoven, *Fuel Process. Technol.* **2012**, *104*, 265.
- [50] C. Liu, A. Marafee, R. Mallinson, L. Lobban, *Appl. Catal. A Gen.* **1997**, *164*, 21.
- [51] O. Mutaf-yardimci, A. Saveliev, A. A. Fridman, L. A. Kennedy, *Int. J. Hydrogen Energy* **1998**, *23*, 1109.
- [52] S. L. Yao, E. Suzuki, N. Meng, A. Nakayama, *Plasma Chem. Plasma Process.* **2002**, *22*, 225.
- [53] B. Zhu, X. Li, J. Liu, A. Zhu, *Int. J. Hydrogen Energy* **2012**, *37*, 16916.
- [54] H. Lee, H. Sekiguchi, *J. Phys. D. Appl. Phys.* **2011**, *44*, 274008.
- [55] A. Indarto, J. Choi, H. Lee, H. Song, *Energy* **2006**, *31*, 2986.
- [56] M. J. Gallagher, R. Geiger, A. Polevich, A. Rabinovich, A. Gutsol, A. Fridman, *Fuel* **2010**, *89*, 1187.
- [57] A. Indarto, J. Choi, H. Lee, H. K. Song, *J. Nat. Gas Chem.* **2005**, *14*, 13.
- [58] X. Guofeng, D. Xinwei, *Energy* **2012**, *47*, 333.
- [59] C. S. Kalra, A. F. Gutsol, A. A. Fridman, *IEEE Trans. Plasma Sci.* **2005**, *33*, 32.
- [60] X. Tu, J. C. Whitehead, *Int. J. Hydrogen Energy* **2014**, *39*, 9658.
- [61] Z. Bo, J. Yan, X. Li, Y. Chi, K. Cen, *Int. J. Hydrogen Energy* **2008**, *33*, 5545.
- [62] A. E. E. Putra, S. Nomura, S. Mukasa, H. Toyota, *Int. J. Hydrogen Energy* **2012**, *37*, 16000.
- [63] J. Yuan, X. Zhong, S. Tan, *J. Nat. Gas Chem.* **2010**, *19*, 605.
- [64] B. Gaudernack, S. Lynam, *Int. J. Hydrogen Energy* **1998**, *23*, 1087.
- [65] H. Lesueur, A. Czernichowski, J. Chapelle, *Int. J. Hydrogen Energy* **1994**, *19*, 139.
- [66] M. Steinberg, H. C. Cheng, *Int. J. Hydrogen Energy* **1989**, *14*, 797.
- [67] C. B. Laflamme, J. W. Jurewicz, D. V. Gravelle, M. I. Boulos, *Chem. Eng. Sci.* **1990**, *45*, 2483.
- [68] V. Goldfarb, R. Tiwary, J. Woodroffe, *Plasma Chem. Plasma Process.* **1989**, *9*, 119S.
- [69] J. R. Fincke, R. P. Anderson, T. Hyde, B. A. Detering, R. Wright, R. L. Bewley, D. C. Haggard, W. D. Swank, *Plasma Chem. Plasma Process.* **2002**, *22*, 107.
- [70] A. Czernichowski, P. Czernichowski, A. Ranaivosoloarimanana, in Proc. 11th World Hydrog. Energy Conf., Stuttgart, Germany, **1996**, pp. 637.
- [71] A. I. Barbaritskii, M. A. Deminskii, S. A. Demkin, V. K. Zhivotov, *High Energy Chem.* **1999**, *33*, 45.
- [72] X. Li, C. Shi, Y. Xu, K. J. Wang, A. M. Zhu, *Green Chem.* **2007**, *9*, 647.
- [73] E. C. Neyts, K. Ostrikov, M. K. Sunkara, A. Bogaerts, *Chem. Rev.* **2015**, *115*, 13408.
- [74] E. C. Neyts, A. Bogaerts, *J. Phys. D. Appl. Phys.* **2014**, *47*, 224010.
- [75] I. Belov, S. Paulussen, A. Bogaerts, *Plasma Sources Sci. Technol.* **2016**, *25*, 15023.
- [76] N. Amin, *Fuel* **2006**, *85*, 577.
- [77] A. A. Fridman, L. A. Kennedy, *Plasma Physics and Engineering*. Taylor & Francis Routledge, New York **2004**.
- [78] P. Jorgensen, J. Chapelle, A. Czernichowski, K. Meguernes, in 8th Int. Symp. Plasma Chem., Tokyo, Japan, **1987**, 695.
- [79] J. Heo, J.-W. Choi, H. Lee, H. Sekiguchi, H. K. Song, in 18th Int. Symp. Plasma Chem., Tokyo, Japan, **2007**, 554.
- [80] D. H. Lee, K. T. Kim, M. S. Cha, Y. H. Song, *Proc. Combust. Inst.* **2007**, *31 II*, 3343.
- [81] T. Nozaki, A. Hattori, K. Okazaki, *Catal. Today* **2004**, *98*, 607.
- [82] A. Indarto, D. R. Yang, J. Palgunadi, J.-W. Choi, H. Lee, H. K. Song, *Chem. Eng. Process. Process Intensif.* **2008**, *47*, 780.
- [83] H. Long, S. Shang, X. Tao, Y. Yin, X. Dai, *Int. J. Hydrogen Energy* **2008**, *33*, 5510.
- [84] J. Q. Zhang, Y. J. Yang, J. S. Zhang, Q. Liu, *Acta Chim. Sin.* **2002**, *60*, 1973.
- [85] A. M. Ghorbanzadeh, R. Lotfalipour, S. Rezaei, *Int. J. Hydrogen Energy* **2009**, *34*, 293.
- [86] S. L. Yao, F. Ouyang, A. Nakayama, E. Suzuki, M. Okumoto, A. Mizuno, *Energy Fuels* **2000**, *14*, 910.
- [87] M. Li, G. Xu, Y. Tian, L. Chen, H. Fu, *J. Phys. Chem. A* **2004**, *108*, 1687.
- [88] T. Jiang, M. Li, Y. Li, G. Xu, C. Liu, B. Eliasson, *J. Tianjin Univ.* **2002**, *35*, 19.
- [89] C.-J. Liu, R. G. Mallinson, L. L. Lobban, *Appl. Catal. A Gen.* **1999**, *178*, 17.
- [90] H. K. Song, J.-W. Choi, S. H. Yue, H. Lee, B.-K. Na, *Catal. Today* **2004**, *89*, 27.
- [91] L. Bromberg, D. R. Cohn, A. Rabinovich, N. Alexeev, *Int. J. Hydrogen Energy* **1999**, *24*, 1131.
- [92] Y. F. Wang, C. H. Tsai, W. Y. Chang, Y. M. Kuo, *Int. J. Hydrogen Energy* **2010**, *35*, 135.
- [93] H. Le, L. L. Lobban, R. G. Mallinson, *Catal. Today* **2004**, *89*, 15.
- [94] T. Hammer, T. Kappes, M. Baldauf, *Catal. Today* **2004**, *89*, 5.
- [95] T. Nozaki, N. Muto, S. Kado, K. Okazaki, *Catal. Today* **2004**, *89*, 57.
- [96] P. C. Kong, Method for cracking hydrocarbon compositions using a submerged reactive plasma system, 1997, 5,626, 726.
- [97] L. C. Casta, J. A. D. Mu, J. Ancheyta, *Catal. Today* **2014**, *220–222*, 248.
- [98] M. Gharibi, A. Khosravi, M. R. Khani, S. S. Shahabi, E. Dejbani, B. Shokri, *J. Electrostat.* **2015**, *76*, 178.
- [99] H. Hao, B. S. Wu, J. Yang, Q. Guo, Y. Yang, Y. W. Li, *Fuel* **2014**, *149*, 162.
- [100] A. Jahanmiri, M. R. Rahimpour, M. M. Shirazi, N. Hooshmand, H. Taghvaei, *Chem. Eng. J.* **2012**, *191*, 416.
- [101] C. Liu, I. Chernets, H.-F. Ji, J. Smith, A. Rabinovich, D. Dobrynin, A. Fridman, *IEEE Trans. Plasma Sci.* (under review), arXiv preprint arXiv:1602.06036, **2016**.

- [102] A. Rabinovich, A. Fridman, D. Dobrynin, in *22nd Int. Symp. Plasma Chem.* **2015**, IN-19.
- [103] C. Liu, A. Fridman, A. Rabinovich, D. Dobrynin, in *22nd Int. Symp. Plasma Chem.* **2015**, p. O-19.

## AUTHORS' BIOGRAPHIES



**R. Snoeckx**, from the University of Antwerp in Belgium, obtained a master's degree in both environmental science (2010) and chemistry (2012), and is currently finalising his PhD. Combining these two backgrounds he started his PhD work in chemistry at the research group PLASMANT on the

topic of plasma based conversion of greenhouse gases into value-added chemicals and fuels. For this research he actively collaborated with several (inter)national research institutions, for example during research visits at the Instituto Superior Técnico, the King Abdullah University of Science and Technology, and the Drexel Plasma Institute.



**A. Rabinovich** is Associate Director and Research Professor at the A.J. Drexel Plasma Institute. He also is a Director of the Fuel Conversion & Hydrogen Production Lab at the A.J. Drexel Plasma Institute. His research interests include development of new plasma processes and systems for

waste treatment and production of hydrogen rich gas from a wide range of hydrocarbon fuels including diesel, gasoline, JP8, and biofuels. Another special area of interest is plasma assisted direct liquefaction of natural gas with production of valuable gaseous and liquid products.



**D. Dobrynin** has received the MS degree in physical electronics from the Petrozavodsk State University, Petrozavodsk, Russia, in 2008 and the PhD degree from Drexel University, Philadelphia, USA, in 2011. He joined the AJ Drexel Plasma Institute as a research faculty and a director of Applied

Physics Laboratory in 2011. His research interests include experimental plasma and gas discharge diagnostics, discharges in liquid phase, plasma applications in medicine and biology, and plasma engineering.



**A. Bogaerts** was born in 1971. She obtained her PhD degree in chemistry in 1996, from the University of Antwerp in Belgium. After some postdoc years, she became a professor of physical chemistry in 2003, at the same university, and is full professor since 2012. She is head of the

interdisciplinary research group PLASMANT. Her research activities include modeling of plasma chemistry, plasma reactor design, and plasma-surface interactions, mainly for environmental and medical applications. In recent years, special attention is given to CO<sub>2</sub> conversion by plasma and plasma catalysis, as well as plasma for cancer treatment, and for both topics, the group is now also performing experimental research.



**A. Fridman** is Nyheim Chair Professor of the Drexel University, Philadelphia, USA, and director of the A.J. Drexel Plasma Institute. He develops novel plasma approaches to material treatment, fuel conversion, hydrogen production, aerospace engineering, biology, and environmental control.

Dr. Fridman worked and taught as a Professor and researcher in different National Laboratories and Universities of United States, France, and Russia. He published eight books and more than thousand scientific papers, chaired several international plasma conferences and received numerous awards, among others he received the State Prize of the USSR for discovery of selective stimulation of chemical processes in non-thermal plasma together with the Nobel Prize laureate N. G. Basov.

**How to cite this article:** Snoeckx R, Rabinovich A, Dobrynin D, Bogaerts A, Fridman A. Plasma-based liquefaction of methane: The road from hydrogen production to direct methane liquefaction. 2017;14:e1600115. <https://doi.org/10.1002/ppap.1600115>